



## HEXAVALENT CHROMIUM REMOVAL FROM POLLUTED WATER USING NEWLY PREPARED ADSORBENT FROM BARK OF *TAMARINDUS INDICA*

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### ABSTRACT:

Environment is deteriorating day to day due to the industrial pollution, toxic chemicals etc. which lead to the accumulation of contaminants inclusive of heavy metals in waste water. The water polluted by industrial effluents disturbs its normal use for agriculture and aquatic life. Due to the toxicity, non-biodegradability and persistent nature, removal of toxic heavy metals become an absolute necessity. Hexavalent chromium metal ion is one of the major pollutant in the environment which is frequently present in waste water released from various industrial units. In the recent research paper reports, the potential application of self prepared activated carbon derived from the bark of *Tamarindus indica* for removal of hexavalent chromium from aqueous solution. The batch experiments were carried out to investigate the effect of significant process parameters such as pH, contact time, adsorbent dosage and initial Cr(VI) ion concentration. The maximum adsorption of Cr(VI) on bio-sorbent was found at pH 4.5. The removal of hexavalent chromium from aqueous solution increases with increase in contact time and adsorbents dose. The maximum removal capacity for Cr(VI) have been noticed to be 93.70%. This investigation verifies that the activated carbon derived from *Tamarindus indica* bark can be used as a valuable adsorbent material for removal of hexavalent chromium from contaminated water and may contribute in pollution control.

**Keywords:** - Activated Carbon, Adsorption, *Tamarindus indica* (TI), Hexavalent chromium.

### INTRODUCTION :

The pollution of water resources due to the indiscriminate disposal of heavy metals has been causing world wise concern for the last few decades. The removal of heavy metal contaminants from aqueous solutions is one of the most important environmental concerns because metals are biorefractory and are toxic to many life forms.<sup>1,2</sup> Metals which are significantly toxic to human beings and ecological environments, include chromium, copper, lead, mercury, cadmium, nickel, iron etc.<sup>3</sup> This problem has received considerable attention in recent years. These heavy metals are toxic to aquatic flora and fauna even in relatively low concentrations. Some of these are capable of being assimilated, stored and concentrated by organisms.<sup>4,5</sup> Generally, metals that form compounds with biological constituents can be

toxic, carcinogenic or mutagenic even in a very low concentration.<sup>6</sup> Unlike organic pollutants; the majority of which are susceptible to biological degradation; metal ions do not degrade into harmless end products.<sup>7</sup> They are stable and persistent environmental contaminants since they cannot be degraded or destroyed. Modernization, industrialization and urbanization has drastically changed the biogeochemical cycle and balance of some metals in the ecosystem.

Chromium(VI) is one of the most strategic and critical material having wide range uses in various industries. Chromium compounds present in the effluents as a result of electroplating, metal finishing, magnetic tapes, wood preservation, leather tanning, pigments, and chemical manufacturing industries<sup>8,9</sup>. The maximum concentration limit for chromium

discharge into inland surface water is 0.1mg/l and it should not exceed to 0.05mg/l in potable water<sup>10</sup>. Chromium is a toxic metal of widespread use in many industries such as electroplating, leather tanning, metal finishing, textile industries and chromate preparation.<sup>11,12</sup> Various treatment technologies have been developed for the purification of water and waste water contaminated by heavy metals. The most commonly used methods for the removal of metals ions from waste water are chemical precipitation, solvents extraction, oxidation, reduction, electrolytic extraction, reverse osmosis, ion-exchange, adsorption etc<sup>13</sup>. Most conventional techniques are therefore not profitable as industrial-scale method for removal of hexavalent chromium from aqueous medium<sup>14</sup>. Amongst all these methods adsorption is highly effective and economical. Adsorbents derived from various kinds of biomass, mainly industrial and agricultural by-products, such as clay, corncob, and fly ash <sup>15,16,17</sup> which are used for the removal of various contaminants from the water environment. Among them, a large number of low-cost biosorbents developed from different by-products, such as green coconut shell <sup>18</sup>, sugarcane bagasse <sup>19</sup>, coffee husk<sup>20</sup>, rice husk <sup>21</sup>, mango kernel <sup>22</sup>, maize cob <sup>23</sup>, sawdust <sup>24</sup>, hazelnut shell <sup>25</sup>, groundnut hull <sup>26</sup>, sugarcane bagasse <sup>27</sup>, pea pod peel <sup>28</sup>, avocado seed kernel <sup>29</sup>, tea waste <sup>30</sup> and olive bagasse <sup>31</sup> had been used for aqueous Cr(VI) adsorption. It is a growing need to derive activated carbons from cheaper and locally available waste materials. Several research workers used different low cost adsorbents from agriculture wastes such as coconut coir pith, sawdust, rice husk, banana pith, cottonseed hulls, apples wastes, sugarcane bagasse, peanut hull etc. for the removal of Cr(VI) from water and waste water. The present work reports the studies carried out for the removal of Cr(VI) from aqueous solution using activated carbon derived

from bark of *Tamarindusindica*. The newly prepared adsorbent was characterized by FTIR and Scanning Electron Microscopy (SEM) studies. Batch isothermal equilibrium method was conducted at 310K to evaluate the efficiency of newly synthesized adsorbent for removal of Cr(VI) from the aqueous solution. Experiments were carried out to study the effect of pH, adsorbent dosage, contact time and initial Cr(VI) concentration. The newly synthesized composite have been proved to be very good adsorbent which can be successfully used for removal of carcinogenic hexavalent chromium from aqueous solution.

## MATERIALS AND METHODS :

### Chemicals

The chemicals used in the investigation were analytical grade and procured from Merck (Mumbai, India).

### Preparation of Activated Carbon from the bark of *Tamarindusindica*(TIAC)

The bark of *Tamarindusindica* was collected from the local area. The bark was cut into small pieces, washed with tap water to remove the sand particles and then treated with formaldehyde to avoid release its colour into aqueous solution. It was washed several times with deionized water and sun dried for 6 days. After drying, the bark was subjected to pyrolysis process for carbonization using Muffle Furnace at 800-900°C for 7-8 hrs so that volatile constituents were removed and residue was converted into a char. The char was then subjected to microwave activation in microwave oven. The input power of microwave equipment was set at 360 W for 30 min. The resulting activated carbon particles were ground and sieved in 120-200 mm size. This activated carbon was then washed with double distilled water and dried at 105°C for 3 hrs and stored in air tight bottle.

### Characterization of TIAC

Characterization of TIAC was done by FTIR (Fig.1) and SEM (Fig.2)

### Adsorption Studies

Working solution were prepared by progressive dilution of stock solution of Cr(VI). Removal of Cr(VI) using TIAC was carried out by batch equilibrium method. The influence of various parameters such as effect of pH, contact time, adsorbent dosage and initial Cr(VI) ion concentration were studied at 30°C, taking 25mg/l of initial Cr(VI) concentration and 5g/l of adsorbent dose. The effect of adsorbent dosage was studied by varying them from 1 to 10g/l. The effect of initial Cr(VI) concentration was studied by changing concentration from 10 to 100mg/l with fixed adsorbent dose of 5g/l at 30°C. The residual concentration of Cr(VI) were measured by Atomic Absorption Spectrophotometer (AAS) using air acetylene flame.

## RESULTS AND DISCUSSION :

### Characterization of TIAC

**Fig.1** represents FTIR of TIAC. The absorption band at  $3438.57\text{cm}^{-1}$  correspond to -OH stretching vibration of water, hydroxyl group and -NH stretching of free amino groups. The signals at  $1612.03$  and  $2284.47\text{cm}^{-1}$  (weak) are assigned to acylamino group and S-H vibrations respectively. The band at  $1314.26\text{cm}^{-1}$  corresponds to C-H and O-H deformation vibrations. Absorption band at  $1230\text{cm}^{-1}$  could be attributed to C-OH stretching. A peak at  $779.13\text{cm}^{-1}$  is indicative of -CH<sub>2</sub> rocking.

The surface morphology of TIAC has been characterized by SEM (**Fig.2**) which indicates porous and fibrous structure of the material under investigation. The shapes of the TIAC particle were nearly spherical and their size ranges from 5 to 10  $\mu\text{m}$ . The surface is also associated with cavity like structure. The surface was rough, uneven, gaped and folded due to the drying procedure and cavity. This is requisite for an effective adsorbent.

### Effect of pH

The effect of pH on the adsorption of chromium(VI) by TIAC was studied at pH 1 to 8. From Fig.3, it is clear that the removal of Cr(VI) increases with an increase in pH from 1.0 to 4.5 and it is optimum at 4.5. The percentage of adsorption increases from 60 to 96.45% as pH was increased from 1 to 4.5. The percentage of adsorption decreases steadily to 93% when pH was increased above 4.5 and it was further decrease to 75% as pH was raised to 8.

### Effect of Contact Time

Adsorption experiments were conducted as a function of contact time and results have shown in fig.4. It can be observed that Cr(VI) removal ability of TIAC increased with increase in contact time before equilibrium was reached. Other parameters such as dose of CCTIAC, pH of solution and initial concentration were kept optimum. It can be seen from fig.4 that Cr(VI) removal efficiency increased from 58 to 95% when contact time was increased from 60 to 100 min. Optimum contact time for TIAC was found to be 100 min. Cr(VI) removal efficiency remained nearly constant after 100 min i.e. equilibrium time.

### Effect of Adsorbent Dosage

Fig.5 shows the effect of dosage on the removal of Cr(VI) which was studied by varying the amount of TIAC from 0.5 to 10g/l while keeping other parameters (pH, contact time and initial concentration) constant. It is clear from the figure that percentage removal of Cr(VI) increased with the increase in TIAC doses and it was found to be maximum i.e. 93.5% at the dose of 4.0g/l. This is due to availability of more adsorbent (more surface area). It indicates that by increasing the TIAC dosages, the adsorption efficiency for Cr(VI) removal increases. After 4g/l dose of TIAC, the adsorption efficiency remained constant because the maximum adsorption set in and amount of Cr(VI) ion present in the

solution bounded to adsorbent remains nearly constant after this dose.

#### Effect of initial metal ion concentration

The effect of initial metal ion concentration on the percentage removal of hexavalent chromium by TIAC has shown in fig.6. It can be seen that the percentage removal of Cr(VI) decreases with the increase in initial Cr(VI) concentration. The adsorbent dose was maintained 5g/l. The result shows the decrease in removal from 93 to 60%. This can be justified by the fact that adsorbent have limited number of active sites which are saturated beyond certain concentration of adsorbate.

#### CONCLUSION :

- The activated carbon was derived successfully from the bark of *Tamarindusindica* and characterized employing FTIR and SEM studies.
- The newly developed modified activated carbon has high porous structure and excellent surface area.
- TIAC was most effective for Cr(VI) removal. At pH 4.5, 96% of Cr(VI) was removed from aqueous solution. Adsorption was found to pH dependent. Above pH 4.5, decrease in Cr(VI) removal was noticed.
- The increase in percent removal capacity for Cr(VI) was observed with increased of adsorbent doses and contact time. Maximum removal is 96% for 4.0 g/l dose and 110 min. of contact time.
- The newly obtained activated carbon under present investigation can be successfully employed for Cr(VI) abatement from contaminated water and thus can be used for water/ wastewater treatment.

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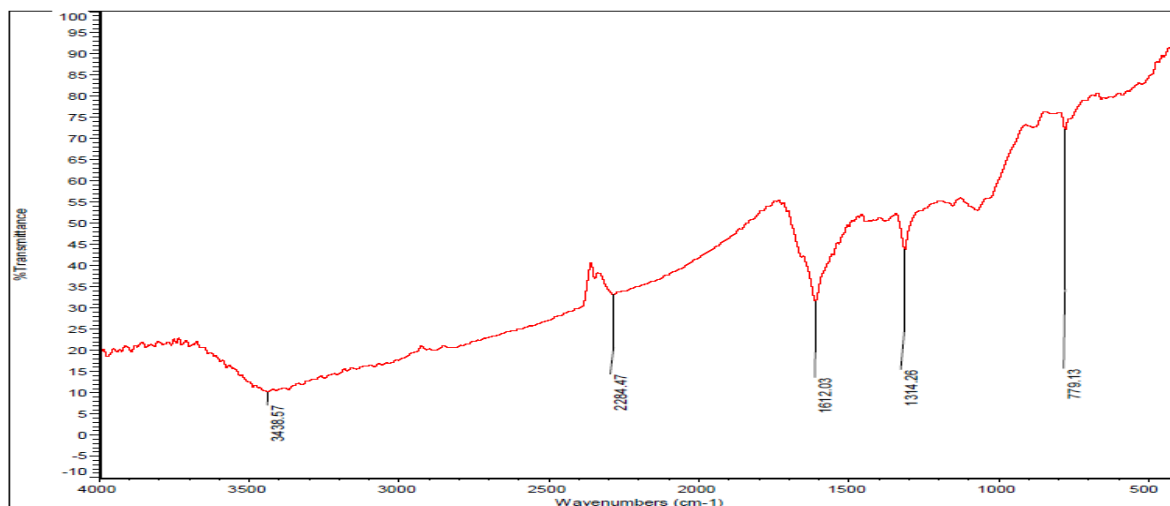
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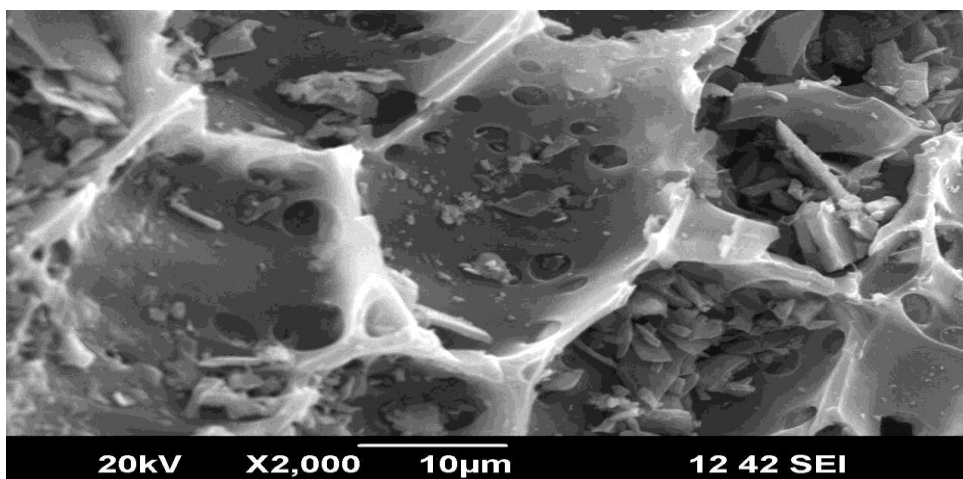
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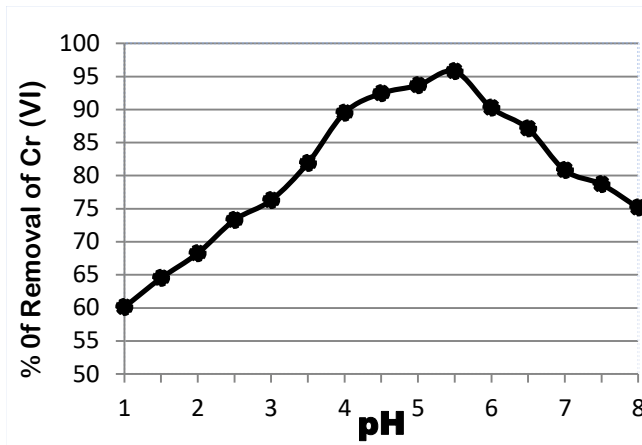
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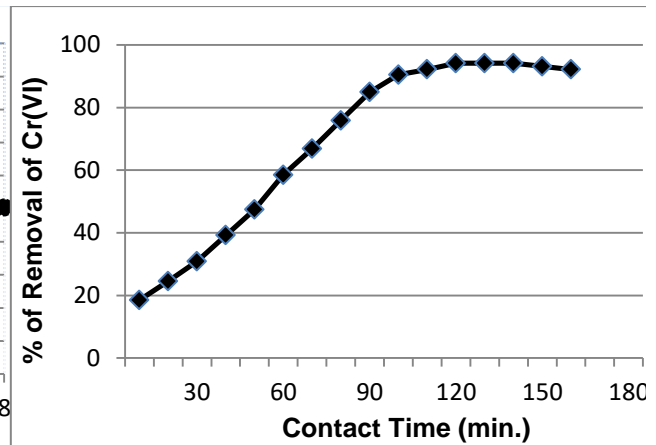
**Fig. 1: FT-IR Spectrum of activated carbon of *Tamarindus Indica*(TIAC)**



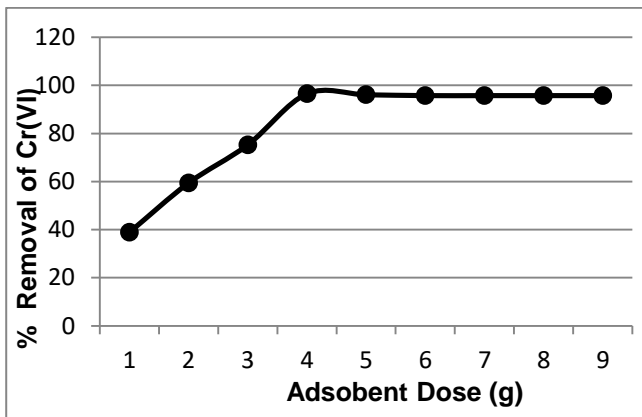
**Fig. 2 : Scanning electron micrographs (SEM) of the TIAC at 2000 x**



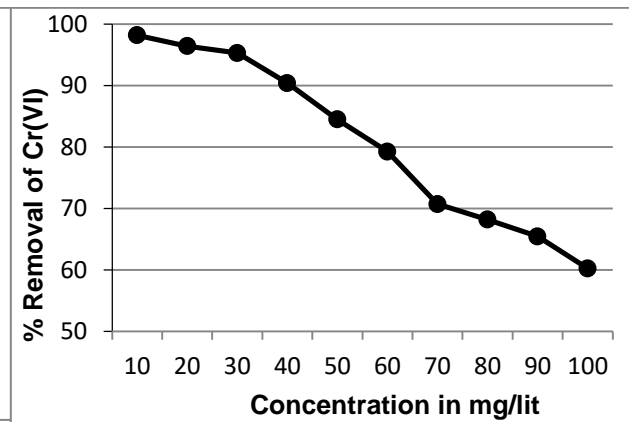
**Fig.3. Effect of pH on Cr(VI) adsorption**



**Fig.4. Effect of Contact time on Cr(VI) adsorption**



**Fig.5. Effect of Adsorbent doses on Cr(VI) adsorption**



**Fig.6. Effect of concentration on Cr(VI) adsorption**